Electrolysis of Aluminium oxide Al2O3

Before adding to the cell, Al2O3 is mixed with to lower its .

A layer of around the inside of the cell forms the negative electrode. The positive electrodes are also made from graphite.

In the cell, the mixture of Al2O3 and cryolite is and a is passed.

Molten Al is formed at the electrode, where the Al3+ ions are to Al atoms. The liquid Al runs to the bottom of the cell, where it is tapped off.

At the positive electrode, the O2- ions are and form gas.

The graphite in the positive electrodes with oxygen, producing and so the electrodes must be replaced regularly.

* Half equation at negative electrode:
* Half equation at positive electrode:

Keywords:

**graphite oxidised oxygen react carbon dioxide cryolite**

**melting point current negative reduced melted**